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asked for number of moles and number of molecules strategy a use the molecular formula of the compound to calculate its molecular mass in grams per mole b convert from mass to moles by dividing the mass given by the compound's molar mass c convert from moles to molecules by multiplying the number of moles by avogadro's number solution think of mole as a descriptive unit of measure like a dozen a dozen is always 12 of something compare a dozen eggs versus a dozen textbooks for example in the same way the mole always refers to 6.022×10^{23} of something so a mole of carbon versus a mole of iron still means the same number of carbon atoms and iron atoms you might need calculator using the information in the table calculate the number of moles in a 2.03 kg sample of citric acid c a 6 h a 8 o a 7 write your answer using three significant figures mol c a 6 h a 8 o a 7 the mole symbol mol is the si unit of amount of substance one mole of a substance contains the same number of the stated particles atoms molecules or ions as one mole of any other substance one mole contains 6.02×10^{23} particles e g atoms ions molecules this number is known as the avogadro constant for example 6 h 6 1 6 30 g find the mass of 3.65×10^{-2} mol K_2SO_4 3.65×10^{-2} mol 174 g K_2SO_4 1 mole 6 351 g K_2SO_4 how many representative particles are in 2.5 mol H_2O_2 2.5 mol 6.02×10^{23} particles 1 mole 1.51×10^{24} particles H_2O_2 chem learn with flashcards games and more for free definition physics and chemistry the simplest structural unit of an element or compound molar mass hydrate avogadro's number mole 3 of 19 definition the mass in grams of 1 mol of a substance molecular mass molar mass avogadro's number atomic mass 4 of 19 definition 6.022×10^{23} answer 2 78 10 22 molecules how big is a mole it is very large suppose you had a mole of dollar bills that need to be counted if everyone on earth about 6 billion people counted one bill per second it would take about 3.2 million years to count all the bills a mole of sand would fill a cube about 32 km on a side in the field of chemistry a mole is defined as the amount of a substance that contains exactly $6.02214076 \times 10^{23}$ elementary entities of the given substance the number $6.02214076 \times 10^{23}$ is popularly known as the avogadro constant and is often denoted by the symbol n_A a chemistry lead did this video help you the mole higher tier only the mole chemical amounts are measured in moles the symbol for the unit mole is mol one mole of a substance contains the same number of the stated particles atoms molecules or ions as one mole of any other substance the mole is a standard si unit used primarily in chemistry this is a collection of ten chemistry test questions dealing with the mole a periodic table will be useful to complete these questions answers appear after the final question 01 of 11 question 1 how many moles of copper are in 6 000 000 atoms of copper 02 of 11 question 2 act math the ultimate guide to the mole the albert team last updated on march 1 2022 the mole like grams or meters is a unit of measurement that is part of the international system of units si system essentially the mole is used to define a specific number of something similar to using the word dozen to define 12 things chemistry student edition basic answer key chapter 10 the mole 10 1 the mole concept questions 1 imagine that you have 1 mole of coins each of which is 1.5 mm thick if they were placed in a single stack how tall would the stack be in km study with quizlet and memorize flashcards containing terms like avogadro's number n_A avogadro's theory empirical formula and more 1 find the number of atoms of phosphorus p in 3.44 moles of phosphorus 2 what is the mass of 0.38 moles of cobalt co 3 how many moles of nickel ni is 3.88×10^{25} atoms of nickel 4 how many atoms is 3.75 moles of iron fe 5 find the number of moles of sodium na in 145 g of sodium 6 how many moles is 0.55 g of magnesium mg 7 goals for the activities at the completion of this unit students will 1 have a conceptual understanding of the mole as the method of counting items and finding the mass of items that can't be seen 2 be able to calculate the number of items molecules atoms ions and formula units if given the number of moles form 3 chemistry questions and answers on the mole chemistry form 3 questions and video answers on the mole many questions from

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the mole symbol mol is the SI unit of amount of substance one mole of a substance contains the same number of the stated particles atoms molecules or ions as one mole of any other substance one mole contains 6.02×10^{23} particles e.g. atoms ions molecules this number is known as the avogadro constant for example

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sodium 6 how many moles is 0 55 g of magnesium mg 7

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