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rate of diffusion is inversely related to the molecular mass heavier molecules such as hcl here diffuse more slowly than lighter molecules answers a and d are too short answer b too long simple proportion would suggest about 73 s but the rate of diffusion is inversely related to the square

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igsce chemistry answers b the mass of the empirical formula is 4 12 8 1 16 72 since this is equal to the relative formula mass the molecular formula is the same as the empirical formula that is  $\text{C}_4\text{H}_8\text{O}_{10}$  you know the mass of anhydrous sodium sulfate 142 g you can work out the mass of water of crystallisation

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1 place a glass of water outside it will freeze if the temperature is below  $0^\circ\text{C}$  3 a law states a consistently observed phenomenon can be used for prediction b theory a widely accepted explanation of the behavior of matter c hypothesis a tentative explanation can be investigated by experimentation 5

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11 a 133 cs b127 i c 31 p 3 d 57 co 3 13 a carbon 12 12 c b this atom contains six protons and six neutrons there are six electrons in a neutral 12 c atom the net charge of such a neutral atom is zero and the mass number is 12 c the preceding answers are correct

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for the correct expression  $K_h = \frac{[H_2O][HCOO^-]}{[HCOOH]}$  point for the correct calculated concentration of  $[H_2O] = \frac{[HCOOH]}{[HCOO^-]} = \frac{0.25}{0.25} = 1$  let  $[H_2O] = x$  then  $1.8 \times 10^{-4} = \frac{x}{0.25}$  assume  $x = 0.25$  then  $1.8 \times 10^{-4} = 0.25$

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