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excerpt from the behaviour of copper anodes in chloride solutions my electrolyses were consequently carried out under conditions that allowed the concentrations of the various constituents of the solution at the anode to be ascertained and the concentrations so determined were compared with the requirements of the mass law about the publisher forgotten books publishes hundreds of thousands of rare and classic books find more at forgottenbooks.com this book is a reproduction of an important historical work forgotten books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy in rare cases an imperfection in the original such as a blemish or missing page may be replicated in our edition we do however repair the vast majority of imperfections successfully any imperfections that remain are intentionally left to preserve the state of such historical works a process has been developed for the removal of chromium copper iron manganese nickel lead tin and zinc impurities from an acid aqueous solution of americium chloride using a mercury cathode cell operating at 5 10 amperes and 16 18 volts the americium is not affected the process may also be used to remove other impurity elements precipitates produced from mixed AlCl_3 NaCl solutions exhibited an increasing tendency to form pseudoboehmite over bayerite with increasing sodium chloride concentration upon neutralizing mixed AlCl_3 MgCl_2 systems the final solution pH decreased significantly with increasing magnesium chloride concentration furthermore increasing concentrations of MgCl_2 promoted the precipitation of poorly crystalline mixed magnesium aluminum hydroxide hydrates identified as $\text{Mg}_6\text{Al}_2(\text{OH})_{18} \cdot 4\text{H}_2\text{O}$ by xrd increases in temperature up to 125 °C aging time 48 hours and OH/Al ratio 3/45 were unsuccessful in improving the crystallinity or transformation of these precipitates aqueous solutions of ferrous ammonium sulfate with cupric chloride were studied as a means for determining the uniformity of the dose rate around a multikilocurie cylindrical array of cobalt 60 ferric ion was measured spectrophotometrically at 305 millimicrons the solution selected to satisfy the requirements for dosimetry contained 0.0005 M ferrous ammonium sulfate and 0.005 M cupric chloride in 0.001 N sulfuric acid spectrophotometric measurements of ferric ion were made in solutions brought to 0.15 N acid concentration instead of the conventional 0.8 N in order to minimize spurious oxidation the molar extinction coefficient at this normality proved to be the same as that in 0.8 N solutions the g value as compared to the standard ferrous ferric dosimeter was 0.65 using these solutions no detectable variation in dose rate was found at symmetrically equivalent positions at the center of the cobalt 60 source the average dose rate at each end of the cylindrical array proved to be 12 percent lower than that at the center the dose rate was also less uniform at the ends varying by 4 percent from point to point at a distance of fifteen inches from the center of the source array the dose rate decreased as the square of the distance author this two volume set provides a full account of hydrometallurgy filled with illustrations and tables this work covers the flow of source material from the mined or concentrate state to the finished product it also highlights ion exchange carbon adsorption and solvent extraction processes for solution purification and concentration the extensive reference list over 850 makes this set a valuable resource for extraction and process metallurgists researchers and practitioners

The behaviour of copper anodes in chloride solutions 1911

excerpt from the behaviour of copper anodes in chloride solutions my electrolyses were consequently carried out under conditions that allowed the concentrations of the various constituents of the solution at the anode to be ascertained and the concentrations so determined were compared with the requirements of the mass law about the publisher forgotten books publishes hundreds of thousands of rare and classic books find more at forgottenbooks.com this book is a reproduction of an important historical work forgotten books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy in rare cases an imperfection in the original such as a blemish or missing page may be replicated in our edition we do however repair the vast majority of imperfections successfully any imperfections that remain are intentionally left to preserve the state of such historical works

The Behaviour of Copper Anodes in Chloride Solutions (Classic Reprint) 2017-10-29

a process has been developed for the removal of chromium copper iron manganese nickel lead tin and zinc impurities from an acid aqueous solution of americium chloride using a mercury cathode cell operating at 5 10 amperes and 16 18 volts the americium is not affected the process may also be used to remove other impurity elements

The Application of Thermodynamics to Extractive Metallurgy with Chloride Solutions 1984

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Electrowinning of Copper from Chloride Solutions 1975

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Hydrolysis of U(VI) 1963

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