

Free epub Unit 5 kinetic molecular theory and gas laws (PDF)

the five basic tenets of the kinetic molecular theory are as follows a gas is composed of molecules that are separated by average distances that are much greater than the sizes of the molecules themselves the kinetic molecular theory kmt is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter this theory is based on the following five postulates described here the five main postulates of the kmt are as follows 1 the particles in a gas are in constant random motion 2 the combined volume of the particles is negligible 3 the particles exert no forces on one another 4 any collisions between the particles are completely elastic and 5 the average kinetic energy of the particles is the kinetic molecular theory kmt is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter this theory is based on the following five postulates described here about transcript the kinetic molecular theory kmt can be used to explain the macroscopic behavior of ideal gases in this video we ll see how the kmt accounts for the properties of gases as described by the various gas laws boyle s law gay lussac s law charles s law avogadro s law and dalton s law of partial pressures key points kinetic molecular theory states that gas particles are in constant motion and exhibit perfectly elastic collisions kinetic molecular theory can be used to explain both charles s and boyle s laws the average kinetic energy of a collection of gas particles is directly proportional to absolute temperature only key terms the kinetic molecular theory kmt is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter this theory is based on the following five postulates described here the kinetic molecular theory postulates the experimental observations about the behavior of gases discussed so far can be explained with a simple theoretical model known as the kinetic molecular theory this theory is based on the following postulates or assumptions gases are composed of a large number of particles that behave like hard the kinetic molecular theory of gases assumes that ideal gas molecules 1 are constantly moving 2 have negligible volume 3 have negligible intermolecular forces 4 undergo perfectly elastic collisions and 5 have an average kinetic energy proportional to the ideal gas s absolute temperature there are five main assumptions of the kinetic molecular theory there are no attractive or repulsive forces between gas particles the particles of an ideal gas are separated by great distances compared to their size gas particles have negligible or no volume because of how small and spread apart particles are the kinetic molecular theory kmt is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter this theory is based on the following five postulates described here stephen lower simon fraser university learning objectives make sure you thoroughly understand the following essential ideas which are presented below it is especially important that you know the principal assumptions of the kinetic molecular theory the kinetic molecular theory is a theory that explains the states of matter and is based on the idea that matter is composed of tiny particles that are always in motion the theory helps explain observable properties and behaviors of solids liquids and gases the kinetic molecular theory kmt is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter this theory is based on the following five postulates described here chemical kinetics also called reaction kinetics is the branch of physical chemistry connected with study of how fast chemical reactions go this includes studying how different conditions such as temperature pressure or solvent used affect the speed of a reaction 5 5 the kinetic molecular theory a model for gas behavior page id learning objectives to understand the significance of the kinetic molecular theory of gases the laws that describe the behavior of gases were well established long before anyone had developed a coherent model of the properties of gases

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