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chemical equilibrium when a chemical reaction takes place in a container which prevents the entry or escape of any of the substances involved in the reaction the quantities of these components change as some are consumed and others are formed

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in a chemical reaction chemical equilibrium is the state in which both the reactants and products are present in concentrations which have no further tendency to change with time so that there is no observable change in the properties of the system 1

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chemical equilibrium condition in the course of a reversible chemical reaction in which no net change in the amounts of reactants and products occurs a reversible chemical reaction is one in which the products as soon as they are formed react to produce the original reactants

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unit test introduction to equilibrium dynamic equilibrium direction of reversible reactions types of equilibrium constants writing equilibrium constant and reaction quotient expressions the equilibrium constant K calculating equilibrium constant K_p using partial pressures worked examples calculating equilibrium constants

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13 1 chemical equilibrium learning objectives define chemical equilibrium recognize chemical equilibrium as a dynamic process consider the following reaction occurring in a closed container so that no material can go in or out $H_2 + I_2 \rightleftharpoons 2HI$ this is simply the reaction between elemental hydrogen and elemental iodine to make hydrogen iodide

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quiz unit test about this unit this unit explores the how and why of chemical equilibrium learn about reversible reactions the equilibrium constant le châtelier's principle solubility equilibria and more practice what you've learned and study for the

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section 17.1 equilibrium state and equilibrium constant chemical reactions do not go to completion 100% products even those that look like they do reactions instead reach a point called equilibrium after which the amount of reactants and products no longer change with time this is because all reactions are

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reactions reach chemical equilibrium when the rate of the forward reaction equals the rate of the reverse reaction in this lecture we discuss the nature of chemical equilibrium and of the chemical equilibrium constant we start to consider how external factors can push the equilibrium in one direction or the other

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le chatelier's principle states that when a system in chemical equilibrium is disturbed by a change in temperature pressure or concentration the system will shift in equilibrium composition in a way that tends to counteract this change of variable the three ways that affect this change in equilibrium include

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chemical equilibrium refers to the state of a system in which the concentration of the reactant and the concentration of the products do not change with time and the system does not display any further change in properties table of contents types of chemical equilibrium factors affecting chemical equilibrium examples importance problems

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definition equilibrium as it applies to a chemical reaction system state of a reaction mixture at which the forward reaction rate is equal to the reverse reaction rate note that there is no reference to the amounts of reactants necessary to achieve this state

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many chemical reactions do not go to completion but instead attain a state of chemical equilibrium chemical equilibrium a state in which the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant equilibrium is a dynamic process & the conversions of reactants to products and

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in a chemical equilibrium the forward and reverse reactions occur at equal rates and the concentrations of products and reactants remain constant if we run a reaction in a closed system so that the products cannot escape we often find the reaction does not give a 100 yield of products

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