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answer the questions below based on your reading above and on your knowledge of chemistry 1 dry ice is solid carbon dioxide CO_2 s carbon dioxide occurs naturally as a gas in the atmosphere CO_2 g a which form of carbon dioxide has higher entropy b which form of carbon dioxide has higher enthalpy 2

analyzing oxidation reduction reactions

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1 $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ 2 $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$ 3 $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$ 4 $2\text{K}_2\text{Cr}_2\text{O}_7 + 2\text{H}_2\text{O} + 3\text{S} + 4\text{KOH} \rightarrow 2\text{Cr}_2\text{O}_3 + 3\text{SO}_2$ chemistry form ws10 1 2a analyzing oxidation reduction reactions redox and electrochemistry page 2 reaction element oxidized reduced evan p silberstein 2003 5 $2\text{H}_2\text{O} + \text{O}_2 \rightarrow 2\text{H}_2\text{O}_2$ 6 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

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answer the questions below based on your reading above and on your knowledge of chemistry answer questions 4 by referring to the diagram to the right showing an electrochemical cell the metal at electrode a is silver the metal at electrode c is lead the electrolytes at locations b d and e are potassium nitrate silver nitrate 10

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decay they release high speed particles that bang into
other unstable radioactive atoms hastening their decay as
the process proceeds the amount of radioactive material
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o level chemistry 2019 paper 2 answers a1 a d and g b a
and e c c d g e a a2 a all contains 16 protons but have

differing number of neutrons from 13 14 and 15 b all have 6 protons on the outer shell suggesting that all isotope will gain 2 electrons to form anion c i x sulfur y oxygen z zinc ii y x and z

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to answer the questions that follow 1 which of the substances above has the lowest boiling point 2 which of the substances above has a boiling point of 100 c 3 which of the substances above has the highest boiling point 4

which of the substances above has the highest vapor pressure at

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answer the following questions based on your understanding of the equilibria involved 1 for the following strong acids and bases 100 percent ionized or dissociated what are the hydronium and hydroxide ion concentrations a 3.00 $\times 10^{-4}$ M HNO_3 b 2.50 $\times 10^{-2}$ M Ca(OH)_2 c 4.00 $\times 10^{-3}$ M NaOH

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