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nuclear structure the numbers of protons and neutrons within the nuclei of the atoms involved remains unchanged during chemical reactions this chapter will introduce the topic of nuclear chemistry which began with the discovery of radioactivity in 1896 by french physicist antoine becquerel and has become increasingly important during the

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explore the subatomic world of nuclear chemistry dive into the mysterious realm of radioactivity unravel ancient secrets with radiometric dating and witness the power of nuclear fission and fusion from the smallest atom to the largest stars it s a high energy journey into the nucleus of chemistry itself

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in this chapter we examine some properties of the atomic nucleus and the changes that can occur in atomic nuclei nuclear reactions differ from other chemical processes in one critical way in a nuclear reaction the identities of the elements change

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21 2 patterns of nuclear stability protons in the nucleus repel each other because of charge the strong nuclear force keeps the nucleus together it s the nuclear glue comes from both protons and neutrons neutrons provide strong nuclear force without repulsion but too many neutrons also leads to instability

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nuclear chemistry nuclear transformations rutherford in 1919 performed the first nuclear transformation the transmutations are sometimes represented by listing in order the target nucleus the bombarding particle the ejecting particle and the product nucleus the above equation becomes ${}^{14}_2\text{N} + {}^4_2\text{He} \rightarrow {}^{17}_8\text{O} + {}^1_1\text{H}$

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3 5 chapter summary radioactivity is defined as the emission of particles and electromagnetic rays from the nucleus of an unstable atom six types of radiation produced during nuclear decay were presented within this chapter and include alpha α decay which is composed of two protons and two neutrons and has a 2 charge

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nuclear chemistry is the study of reactions that involve changes in nuclear structure the chapter on atoms molecules and ions introduced the basic idea of nuclear structure that the nucleus of an atom is composed of protons and with the exception of ${}^1_1\text{H}$ ${}^1_0\text{n}$ neutrons

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nuclear chemistry is the study of nuclear reactions and their uses and effects on biological systems nucleons are particles in the nucleus p proton n^0 neutron atomic number is the number of p mass number is the number of p n^0 the mass number is the total number of nucleons in the nucleus

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in this chapter we will examine some of the basic concepts of nuclear chemistry and some of the nuclear reactions that are important in our everyday lives radioactivity learning objective define and give examples of the major types of radioactivity

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