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1 a sample of gas has a pressure of 100 0 torr and 27 0 c calculate the pressure if the temperature is changed to 127 c while the volume remains constant 2 a gas initially at stp is changed to 248 k calculate the final pressure of the gas 3 a gas occupies a volume of 50 0 ml at 27 c and 630 mmhg

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mixed gas laws worksheet 1 how many moles of gas occupy 98 l at a pressure of 2 8 atmospheres and a temperature of 292 k 2 if 5 0 moles of o₂ and 3 0 moles of n₂ are placed in a 30 0 l tank at a temperature of 250 c what will the pressure of the resulting mixture of gases be

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6 a container holds three gases oxygen carbon dioxide and helium the partial pressures of the three gases are 2 00 atm 3 00 atm and 4 00 atm respectively what is the total pressure inside the container 9 atm 7 a gas mixture contains hydrogen helium neon and argon the total pressure of the mixture is 93 6 kpa the partial

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44 name that gas background knowledge air is a mixture of gases the main gases in air are nitrogen and oxygen there are also small amounts of other gases including carbon dioxide helium and argon each of these gases has different properties that are useful to us in different ways

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