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dipole moments occur when there is a separation of charge they can occur between two ions in an ionic bond or between atoms in a covalent bond dipole moments arise from differences in electronegativity the larger the difference in electronegativity the larger the dipole moment about transcript dipole dipole forces occur between molecules with permanent dipoles i e polar molecules for molecules of similar size and mass the strength of these forces increases with increasing polarity polar molecules can also induce dipoles in nonpolar molecules resulting in dipole induced dipole forces created by sal khan the quantity $q d$ the magnitude of each charge multiplied by the vector distance between them is a property of the dipole its value as you can see determines the torque that the dipole experiences in the external field it is useful therefore to define this product as the so called dipole moment of the dipole two equal and opposite charges separated by some distance constitute a dipole the product of the charge and distance between them is called the dipole moment dipole moment represents the strength of the dipole the electric field generated by the dipole is proportional to the dipole moment updated on january 13 2020 a dipole is a separation of opposite electrical charges a dipole is quantified by its dipole moment μ a dipole moment is the distance between charges multiplied by the charge the unit of the dipole moment is the debye where 1 debye is 3.34×10^{-30} c m in physics a dipole from greek δίς dis twice and πόλος polos axis 1 2 3 is an electromagnetic phenomenon which occurs in two ways an electric dipole deals with the separation of the positive and negative electric charges found in any electromagnetic system the zero average 23 of the dipole field 13 does not take into account the contribution of the region between the charges where eq 13 is not valid which is directed mostly against the dipole vector 9 a dipole is a substance that incorporates two equal and opposite charges possessing a charge can be thought of as possessing some amount of energy the charge can be electric magnetic or any other type everything has two sides this is the law of nature that has been followed since the beginning of time the electric dipole moment is a measure of the separation of positive and negative electrical charges within a system that is a measure of the system s overall polarity the si unit for electric dipole moment is the coulomb meter c m the debye d is another unit of measurement used in atomic physics and chemistry a dipole moment is the product of the magnitude of the charge and the distance between the centers of the positive and negative charges it is denoted by the greek letter μ mathematically dipole moment μ charge q distance of separation r 6 5k 454k views 6 years ago new ap general chemistry video playlist this chemistry video tutorial provides a basic introduction into dipole dipole forces of attraction a dipole is a dipole definition a pair of separated electric charges or magnetic poles of equal magnitude but of opposite sign or polarity the quantity $q d$ the magnitude of each charge multiplied by the vector distance between them is a property of the dipole its value as you can see determines the torque that the dipole experiences in the external field it is useful therefore to define this product as the so called dipole moment of the dipole $\vec{p} \equiv q \vec{d}$ a molecule with an area of negative charge and an area of positive charge that are separated many molecules such as those of water are electric dipoles meaning that they have a partial positive charge at one end and a partial negative charge at the other fewer examples the dipole moment of a molecule can be calculated by equation 1 $\mu_i = q_i r_i$ where μ is the dipole moment vector q_i is the magnitude of the i th charge and r_i is the vector representing the position of i th charge the dipole moment acts in the direction of the vector quantity an example of a polar molecule is H_2O 1 two electric charges or magnetic poles that have equal

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