

# Free download Chapter 21 physical properties of gases wikispaces (Download Only)

simon fraser university learning objectives make sure you thoroughly understand the following essential ideas state the three major properties of gases that distinguish them from condensed phases of matter define pressure and explain why a gas exerts pressure on the walls of a container learn about the state of matter known as gas its molecular structure and its transport properties explore the kinetic theory of gases and its historical and scientific significance the properties are given below

- 1 compressibility particles of gas have huge intermolecular spaces in the midst of them by the exertion of pressure much of this space can be diminished and the particles are brought closer thus the volume of gas can be hugely reduced this is termed as compressing the gas figure learn how gases are easy to compress expand to fill their containers and occupy far more space than liquids or solids explore the concepts of pressure volume and atmospheric pressure with examples and diagrams learn how gases behave as a form of matter and how to calculate their pressure volume and temperature using ideal gas laws explore the equations and examples of boyle s law charles s law gay lussac s law and dalton s law also called perfect gas related topics gas ideal gas law ideal gas a gas that conforms in physical behaviour to a particular idealized relation between pressure volume and temperature called the ideal or general gas law learn about the characteristics and behavior of gases such as their composition volume pressure and temperature explore the kinetic theory of matter the ideal gas law and the si units of pressure published 16 february 2022 gases will fill a container of any size or shape evenly gases are a state of matter with no fixed shape or volume image credit miragec via getty images jump to these common properties relate to a unique characteristic of gases gas molecules are incredibly far apart and rarely interact with each other in solids the attractive and repulsive forces between molecules the intermolecular forces are so strong they lock the solid into a fixed shape and size as discussed in our properties of solids module compared to the numbers of molecules involved there are only a few properties of gases that warrant attention here namely pressure density temperature internal energy viscosity heat conductivity and diffusivity
- 6 properties of gases purpose to investigate simple gas laws and explain them in terms of kinetic molecular theory learning outcomes use a molecular level simulation to investigate and interpret phenomena report measurements with associated error estimates measure the temperature and pressure and discover how the properties of the gas vary in relation to each other examine kinetic energy and speed histograms for light and heavy particles explore diffusion and determine how concentration temperature mass and radius affect the rate of diffusion

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