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d Cl_2 2NaBr 2NaCl Br_2 b BaCl_2 K_2CO_3 BaCO_3 2KCl study with quizlet and memorize flashcards containing terms like what coefficients would balance the following equation $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ a 2Al 3O₂ 1Al₂O₃ b 2Al 3O₂ 2Al₂O₃ c 3Al 3O₂ 2Al₂O₃ d 4Al 3O₂ 2Al₂O₃ what coefficients would balance the

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check your answers for the first three formulas using the gizmo AgCl 3 Cu 2 Ag 1 Cl 3 Cu 2 Ba AsO₄ 2 Ba 1 As 2 O 8 NH₄ 3 PO₄ 4 N 3 H 12 P 1 O 4 MnPb 8 Si 2 O 7 3 Mn 1 Pb 8 Si 6 O 21 activity b balancing equations get the gizmo ready erase the chemical formulas in each text box

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which of the following is an example of a double replacement reaction a $3\text{Fe} + 2\text{O}_2 \rightarrow \text{Fe}_3\text{O}_4$ b $\text{BaCl}_2 + \text{K}_2\text{CO}_3 \rightarrow \text{BaCO}_3 + 2\text{KCl}$ c $\text{NiCO}_3 + \text{NiO} \rightarrow \text{CO}_2$ d $\text{Cl}_2 + 2\text{NaBr} + 2\text{NaCl} \rightarrow \text{Br}_2 + \text{BaCl}_2 + \text{K}_2\text{CO}_3 + \text{BaCO}_3 + 2\text{KCl}$ study with quizlet and memorize flashcards containing terms like what coefficients would balance the following equation $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$

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in the gizmo use the up and down arrows to adjust the numbers of hydrogen oxygen and water molecules until the equation is balanced when you are done turn on show summary to check your answer write the balanced equation here $\text{h}_2 \text{ o}_2 \text{ h}_2\text{o}$ solve turn off show summary

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to set up an equation in the chemical equations gizmo type the chemical formulas into the text boxes of the gizmo first type in $\text{h}_2 \text{ o}_2$ in the reactants box and h_2o in the products box this represents the reaction of hydrogen and oxygen gas to form water 1

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in a chemical reaction reactants interact to form products this process is

summarized by a chemical equation in the balancing chemical equations gizmo look at the floating molecules below the initial reaction $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

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summary to check your answer write the balanced equation here $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
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