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lesson 6 concentration of solutions 1 calculate the molarity of each of the following solutions a 2.5 mol of naoh in 0.500 l of solution b 1.8 l of solution containing 3.3 mol kno₃ c 3.0 mol of naoh in 0.500 l of solution d 12 g of hcl in 250 ml of solution 2

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3 matter that is composed of two or more different elements chemically combined in a fixed proportion is classified as 1 a compound 2 an element 3 a mixture 4 a solution 4 a compound differs from an element in that a compound is homogeneous 3 can be decomposed by a chemical reaction has a definite composition 4 has a definite

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questions copy edit show answers see preview 1 multiple choice 1
minute 1 pt a solution is made by dissolving some salt in a beaker of
water the salt is referred to as the solute solution filtrate solvent 2
multiple choice 1 minute 1 pt

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a solution is a homogeneous mixture the major component is the solvent while the minor component is the solute solutions can have any phase for example an alloy is a solid solution solutes are soluble or insoluble meaning they dissolve or do not dissolve in a particular solvent

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unsaturated name 3 matter review 1 draw the following descriptions pure element pure compound mixture of mixture of mixture of two elements an element two diatomic a compound elements a compound 2 classify each of the following with the combination of terms listed below

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arrived in city b at 10 00 then it left city of b at 12 00 and arrived in
city a at 18 00 if the duration of the flight in both directions is the
same what is the duration of the flight in

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