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polar as water part c tested candidates knowledge of how to separate
two soluble salts sodium sulfate and sodium chloride in the solid
state there is no ceiling on achievement all marks on the mark scheme
should be used appropriately all the marks on the mark scheme are
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also be prepared to award zero marks if the candidate's response is
not an electrolytic cell produces 235 g of Al according to the
equation above calculate the number of moles of electrons that must
be transferred in the cell to produce the 235 g of Al 1 point is

earned for the number of moles of Al^{3+} therefore 3 mol e^- transferred per mol Al cereal chemistry may june 2013 volume 90 number 3 cereal chemistry 90 175 180 cereal chemistry 90 181 185 cereal chemistry 90 186 197 the syllabus for h3 chemistry builds on that for h2 chemistry and includes the whole of the h2 chemistry syllabus only content that is not already part of the h2 chemistry syllabus is specifically set out here the h3 chemistry syllabus introduces additional content in two areas namely spectroscopic techniques and further organic mechanisms grade threshold of cambridge igcse chemistry 0620 may june 2013 examination although perovskite solar cells pscs have been developed rapidly the poor quality of perovskite films and difficulty in scalable fabrication in air condition hindered the improvement of psc performance in this work a novel ring molecule 1,4,7,10-tetraazacyclododecane cyclen is introduced to manipulate telephone 01623 467467 fax 01623 450481 email publication orders edexcel.com order code ug036866 summer 2013 for more information on edexcel qualifications please visit our website edexcel.com pearson education limited

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potassium iodide is an ionic solid and dissolves in water a polar solvent it is only partially soluble in ethanol because although ethanol is polar it is not as polar as water part c tested candidates knowledge of how to separate two soluble salts sodium sulfate and sodium chloride in the solid state

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electrolytic cell produces 235 g of al l according to the equation above a calculate the number of moles of electrons that must be transferred in the cell to produce the 235 g of al l 1 point is earned for the number of moles of al al^{3+} 3 e al therefore 3 mol e transferred per mol al

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