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these numerical relationships are known as reaction stoichiometry a term derived from the ancient greek words stoicheion element and metron measure in this article we ll look at how we can use the stoichiometric relationships contained in balanced chemical equations to determine amounts of substances consumed and produced in chemical

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about this unit get ready to better understand chemical reactions with stoichiometry master the art of measuring substances using avogadro s number and explore how the mighty mole helps us predict the outcomes of chemical reactions

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stoichiometry is the calculation of relative quantities of reactants and products in chemical reactions

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stoichiometry limiting reagent example soda fizz comes from sodium bicarbonate and citric acid $\text{H}_3\text{C}_6\text{H}_5\text{O}_7$ reacting to make carbon dioxide sodium citrate $\text{Na}_3\text{C}_6\text{H}_5\text{O}_7$ and water if 1.0 g of sodium bicarbonate and 1.0 g citric acid are reacted which is limiting how much carbon dioxide is produced $3\text{NaHCO}_3 \text{ aq } \text{H}_3\text{C}_6\text{H}_5\text{O}_7 \text{ aq } 3\text{CO}_2$

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stoichiometry page id the quantitative relationship among reactants and products is called stoichiometry the term stoichiometry is derived from two greek words stoicheion meaning element and metron meaning measure on this subject you often are required to calculate quantities of reactants or products

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key concepts to understanding stoichiometry in earlier units you learned about moles molar mass molar volume molarity and avogadro s number stoichiometry is all about using mole ratios and these measurements and manipulating them to get to our desired unit first let s recap some key concepts to understand in stoichiometry

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study guide for content mastery section 12.2 stoichiometric calculations in your textbook read about mole to mole conversion read the following passage and then solve the problems in the equation that follows each problem write in the space provided the mole ratio that can be used to solve the problem

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stoichiometry deals with the quantitative relationships between reactants and products in chemical reactions it involves the use of balanced chemical equations to make predictions about the amounts of substances consumed and produced in a chemical reaction

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stoichiometry in chemistry the determination of the proportions in which elements or compounds react with one another the rules followed in the determination of stoichiometric relationships are based on the laws of conservation of mass and energy and the law of combining weights or volumes

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1 the study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry 2 stoichiometry is based on the law of conservation of mass 3 in any chemical reaction the mass of the products is less than the mass of the reactants 4

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the study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry 2 stoichiometry is based on the law of conservation of mass

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