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answer with a valid justification $\text{CH}_3\text{CH}_2\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{CH}_2\text{COO}^- + \text{H}_3\text{O}^+$ propanoic acid
 $\text{CH}_3\text{CH}_2\text{COOH}$ is a carboxylic acid that reacts with water according to the equation above at 25 °C
the pH of a 50.0 mL sample of 0.20 M $\text{CH}_3\text{CH}_2\text{COOH}$ is 2.79

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