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answer with a valid justification  $\text{CH}_3\text{CH}_2\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{CH}_2\text{COO}^- + \text{H}_3\text{O}^+$  propanoic acid  
 $\text{CH}_3\text{CH}_2\text{COOH}$  is a carboxylic acid that reacts with water according to the equation above at 25 °C  
the pH of a 50.0 mL sample of 0.20 M  $\text{CH}_3\text{CH}_2\text{COOH}$  is 2.79

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