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replace the dry cotton ball with a water soaked cotton ball constant light or temperature or duration of experiment or time of day etc explanation 1 point maximum ensures that glucose is the attractant other variables must be held constant proposed improvement includes but not limited to

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an electrolytic cell produces 235 g of  $\text{Al}$  according to the equation above  
a calculate the number of moles of electrons that must be transferred in the cell to produce the 235 g of  $\text{Al}$  1 point is earned for the number of moles of  $\text{Al}$   $\frac{235 \text{ g}}{27 \text{ g/mol}} = 8.7 \text{ mol}$  therefore 3 mol  $e^-$  transferred per mol  $\text{Al}$

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then as the reaction proceeds at 450 k the total pressure decreases  
explain this decrease in total pressure in terms of what occurs at the  
molecular level it is proposed that the formation of  $\text{C}_2\text{H}_5\text{Cl}$  g proceeds via  
the following two step reaction mechanism step 1  $\text{C}_2\text{H}_4$  g  $\text{HCl}$  g

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